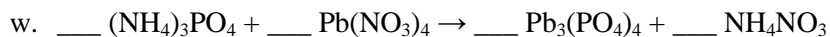
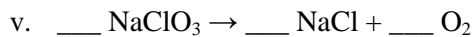
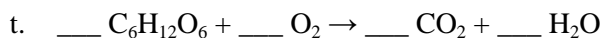
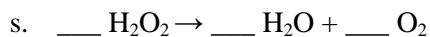
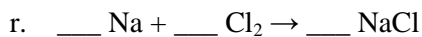
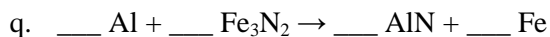
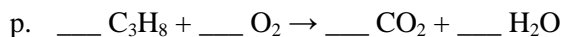
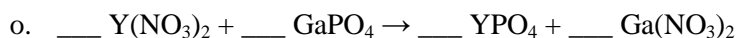
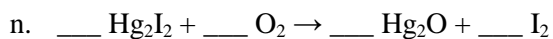
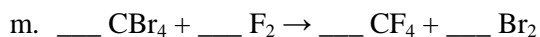
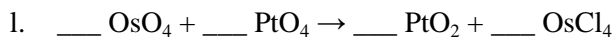
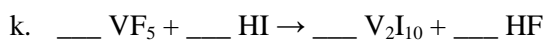
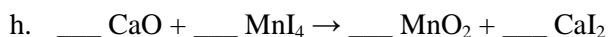
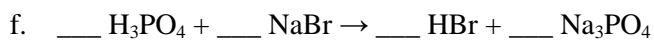
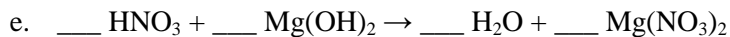
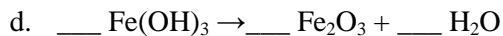
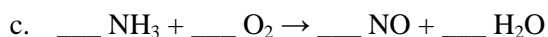
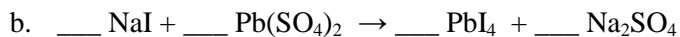
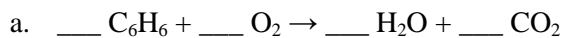


## Balancing Chemical Equations

1. Balance the following equations.



- x.  $\text{___ BF}_3 + \text{___ Li}_2\text{SO}_3 \rightarrow \text{B}_2(\text{SO}_3)_3 + \text{___ LiF}$
- y.  $\text{___ C}_7\text{H}_{17} + \text{___ O}_2 \rightarrow \text{___ CO}_2 + \text{___ H}_2\text{O}$
- z.  $\text{___ CaCO}_3 + \text{H}_3\text{PO}_4 \rightarrow \text{___ Ca}_3(\text{PO}_4)_2 + \text{___ H}_2\text{O} + \text{___ CO}_2$
- aa.  $\text{___ Ag}_2\text{S} \rightarrow \text{___ Ag} + \text{___ S}_8$
- bb.  $\text{___ KBr} + \text{___ Fe}(\text{OH})_3 \rightarrow \text{___ KOH} + \text{___ FeBr}_3$
- cc.  $\text{___ KNO}_3 + \text{___ H}_2\text{CO}_3 \rightarrow \text{___ K}_2\text{CO}_3 + \text{___ HNO}_3$
- dd.  $\text{___ Pb}(\text{OH})_4 + \text{___ Cu}_2\text{O} \rightarrow \text{___ PbO}_2 + \text{___ CuOH}$
- ee.  $\text{___ Cr}(\text{NO}_2)_2 + \text{___ (NH}_4)_2\text{SO}_4 \rightarrow \text{___ CrSO}_4 + \text{___ NH}_4\text{NO}_2$
- ff.  $\text{___ KOH} + \text{___ Co}_3(\text{PO}_4)_2 \rightarrow \text{___ K}_3\text{PO}_4 + \text{___ Co}(\text{OH})_2$
- gg.  $\text{___ Sn}(\text{NO}_2)_2 + \text{___ Pt}_3\text{N}_4 \rightarrow \text{___ Sn}_3\text{N}_2 + \text{___ Pt}(\text{NO}_2)_4$
- hh.  $\text{___ B}_2\text{Br}_6 + \text{___ HNO}_3 \rightarrow \text{___ B}(\text{NO}_3)_3 + \text{___ HBr}$
- ii.  $\text{___ ZnS} + \text{___ AlP} \rightarrow \text{___ Zn}_3\text{P}_2 + \text{___ Al}_2\text{S}_3$