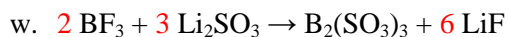
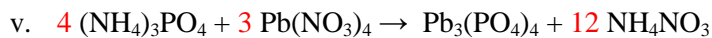
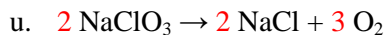
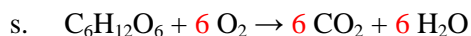
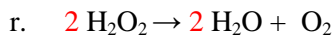
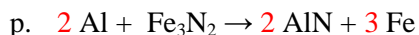
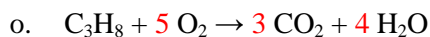
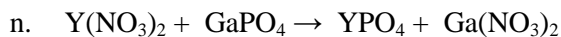
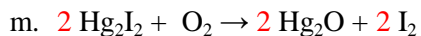
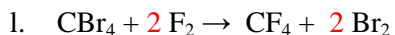
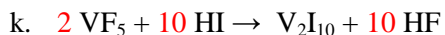
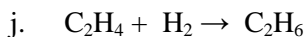
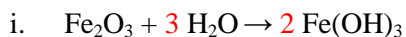
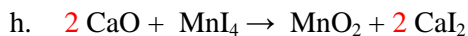
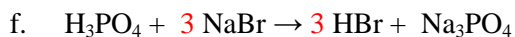
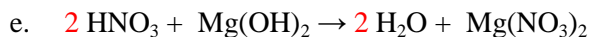
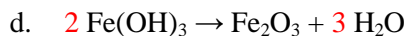
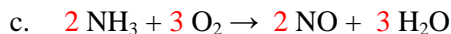
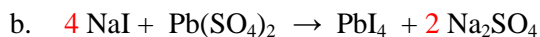
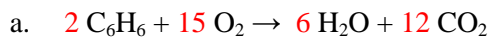


## Balancing Chemical Equations

1. Balance the following equations.



- x.  $4 \text{C}_7\text{H}_{17} + 31 \text{O}_2 \rightarrow 28 \text{CO}_2 + 34 \text{H}_2\text{O}$
- y.  $3 \text{CaCO}_3 + 2 \text{H}_3\text{PO}_4 \rightarrow \text{Ca}_3(\text{PO}_4)_2 + 3 \text{H}_2\text{O} + 3 \text{CO}_2$
- z.  $8 \text{Ag}_2\text{S} \rightarrow 16 \text{Ag} + \text{S}_8$
- aa.  $3 \text{KBr} + \text{Fe}(\text{OH})_3 \rightarrow 3 \text{KOH} + \text{FeBr}_3$
- bb.  $2 \text{KNO}_3 + \text{H}_2\text{CO}_3 \rightarrow \text{K}_2\text{CO}_3 + 2 \text{HNO}_3$
- cc.  $\text{Pb}(\text{OH})_4 + 2 \text{Cu}_2\text{O} \rightarrow \text{PbO}_2 + 4 \text{CuOH}$
- dd.  $\text{Cr}(\text{NO}_2)_2 + (\text{NH}_4)_2\text{SO}_4 \rightarrow \text{CrSO}_4 + 2 \text{NH}_4\text{NO}_2$
- ee.  $6 \text{KOH} + \text{Co}_3(\text{PO}_4)_2 \rightarrow 2 \text{K}_3\text{PO}_4 + 3 \text{Co}(\text{OH})_2$
- ff.  $6 \text{Sn}(\text{NO}_2)_2 + \text{Pt}_3\text{N}_4 \rightarrow 2 \text{Sn}_3\text{N}_2 + 3 \text{Pt}(\text{NO}_2)_4$
- gg.  $\text{B}_2\text{Br}_6 + 6 \text{HNO}_3 \rightarrow 2 \text{B}(\text{NO}_3)_3 + 6 \text{HBr}$
- hh.  $3 \text{ZnS} + 2 \text{AlP} \rightarrow \text{Zn}_3\text{P}_2 + \text{Al}_2\text{S}_3$