So I will do the solution for question 2. Question 1 works the same way.
Question 3 is done exactly like 2b
Question 4 and 5 is like 2c

2 a) Before Titration

$H_2O_{(l)} + HCO_2H_{(aq)} \rightleftharpoons HCOO^-_{(aq)} + H^+_{(aq)}$

I	0.20	C	10^{-7}
C	$-x$	$+x$	$+x$
E	$0.20-x$	x	x

Check
$0.20 - x \cong 0.20$

$K_a = \frac{[HCOO^-][H^+]}{[HCOOH]}$

$1.78\times10^{-4} = \frac{[x][x]}{[0.20]}$

$x = 0.0060$

$pH = -\log[H_3O^+]$
$= -\log[0.0060]$
$= 2.2$

2 b) After 10.0 mL have been added:

$n_{NaOH} = CV$
$= 0.010L \times 0.2\ mol/L$
$= 0.002 mol$

$n_{HCOOH} = CV$
$= 0.025 \times 0.2\ mol/L$
$= 0.005$ moles

$HCOOH + NaOH \longrightarrow H_2O + NaHCOO$
0.005 moles 0.002 moles

becomes

$HCOOH + NaOH \rightleftharpoons H_2O + NaHCOO$
0.003 moles 0.002 moles

new volume = 0.035L
new conc = 0.086 mol/L 0.057 mol/L

$pH = pK_a + \log\frac{[A^-]}{[HA]}$
$= pK_a + \log\frac{[NaHCOO]}{[HCOOH]}$
$= 3.75 + \log\frac{[0.057]}{[0.086]}$
$= 3.57$

2 c) At equivalence: all acid has been converted to salt

$NaOH + HCOOH \rightarrow H_2O + NaHCOO$

0.005 moles in new volume of 50mL
$C = \frac{n}{V} = 0.1\ mol/L$

Will hydrolyse since it is conj of a weak acid

$NaHCOO_{(aq)} + H_2O_{(l)} \rightleftharpoons HCOOH_{(aq)} + OH^-_{(aq)} + Na^+_{(aq)}$

I	0.1		0	10^{-7}
C	$-x$		$+x$	$+x$
E	$0.1-x$		x	x

Check
$0.1 - x \cong 0.1$

since this is a base

$K_b = \frac{K_w}{K_a} = 5.6\times10^{-11}$

$K_b = \frac{[OH^-][HCOOH]}{[NaHCOO]} = \frac{x^2}{0.1} = 2.4\times10^{-6}$

$pOH = -\log[OH^-]$
$pH = 14 - pOH$