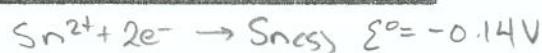
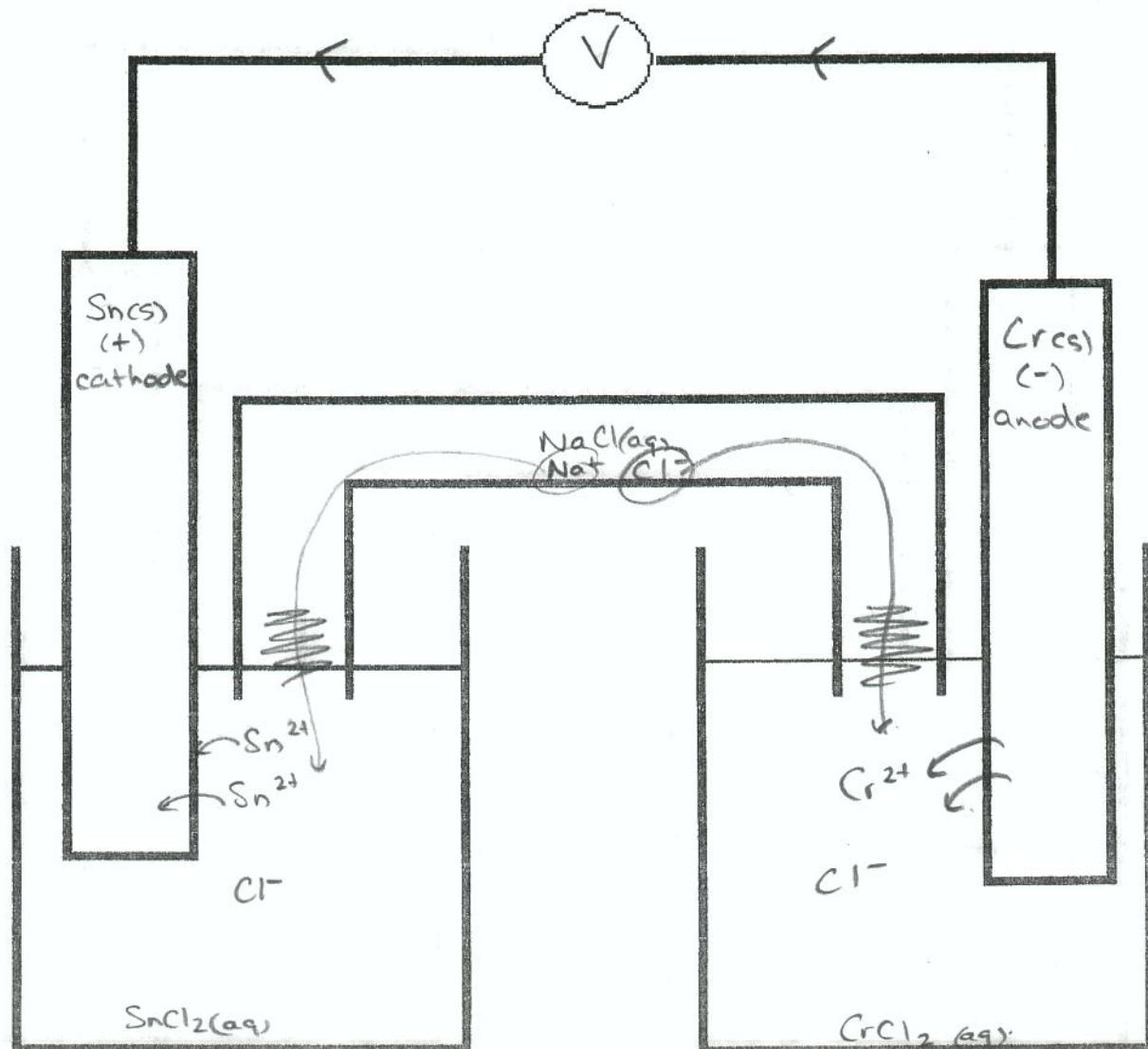
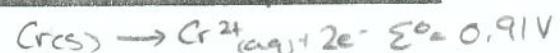


$\text{Sn(s)}$  in  $\text{SnCl}_2\text{(aq)}$  }  $\text{Cr(s)}$  in  $\text{CrCl}_2\text{(aq)}$



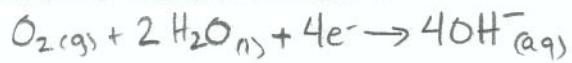
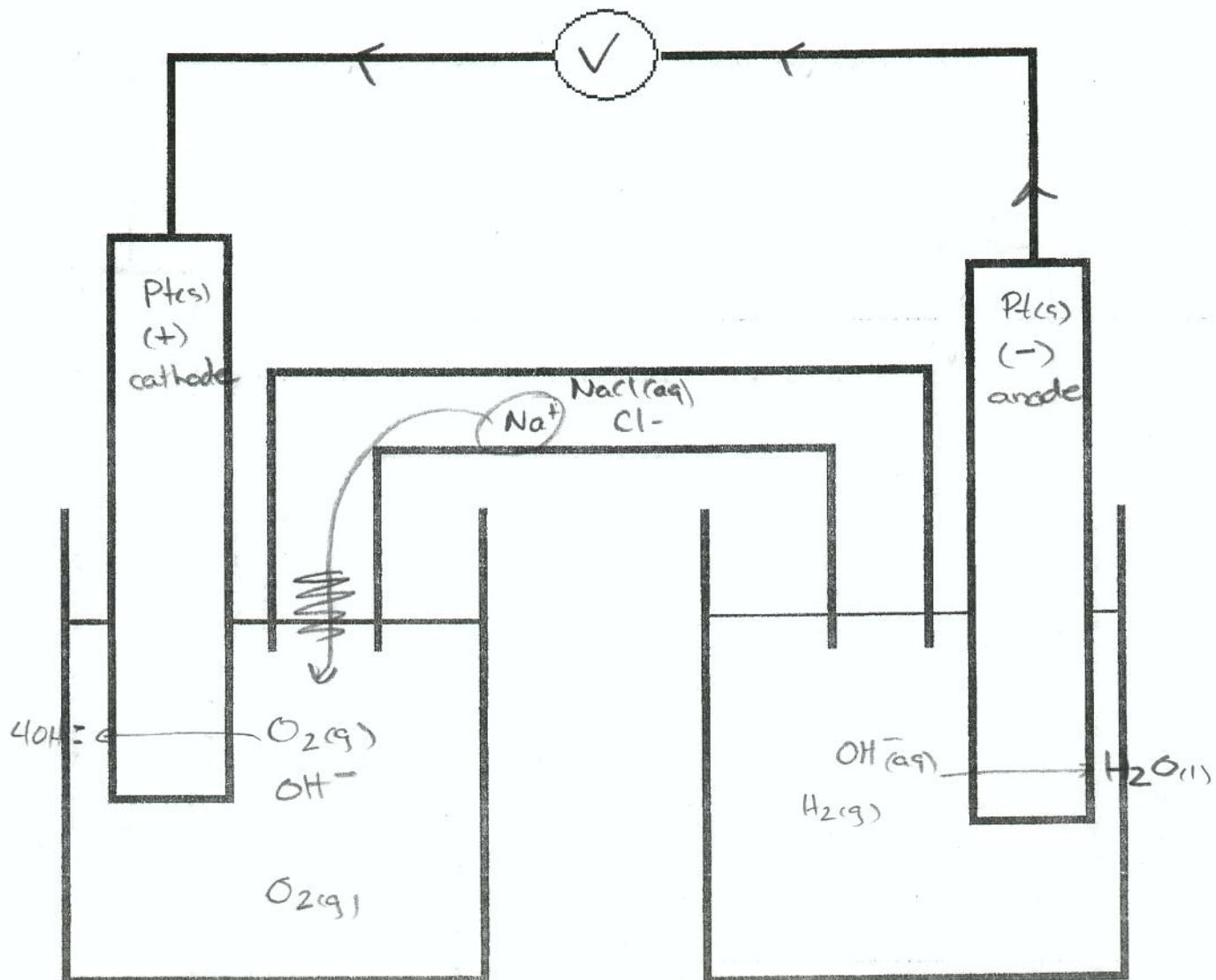
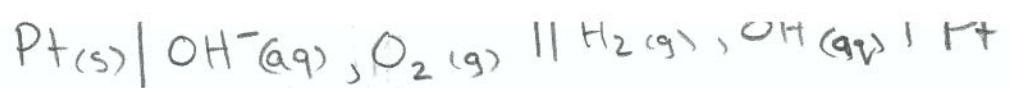
reduction



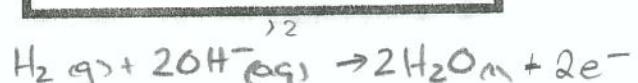
oxidation



$$\begin{aligned} \mathcal{E}^\circ_{\text{cell}} &= \mathcal{E}^\circ_{\text{red}} + \mathcal{E}^\circ_{\text{oxidation}} \\ &= -0.14 \text{ V} + 0.91 \text{ V} \\ &= 0.77 \text{ V} \end{aligned}$$



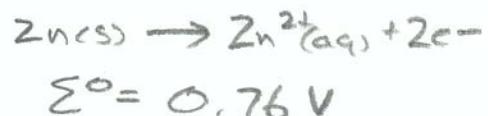
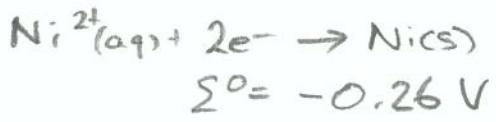
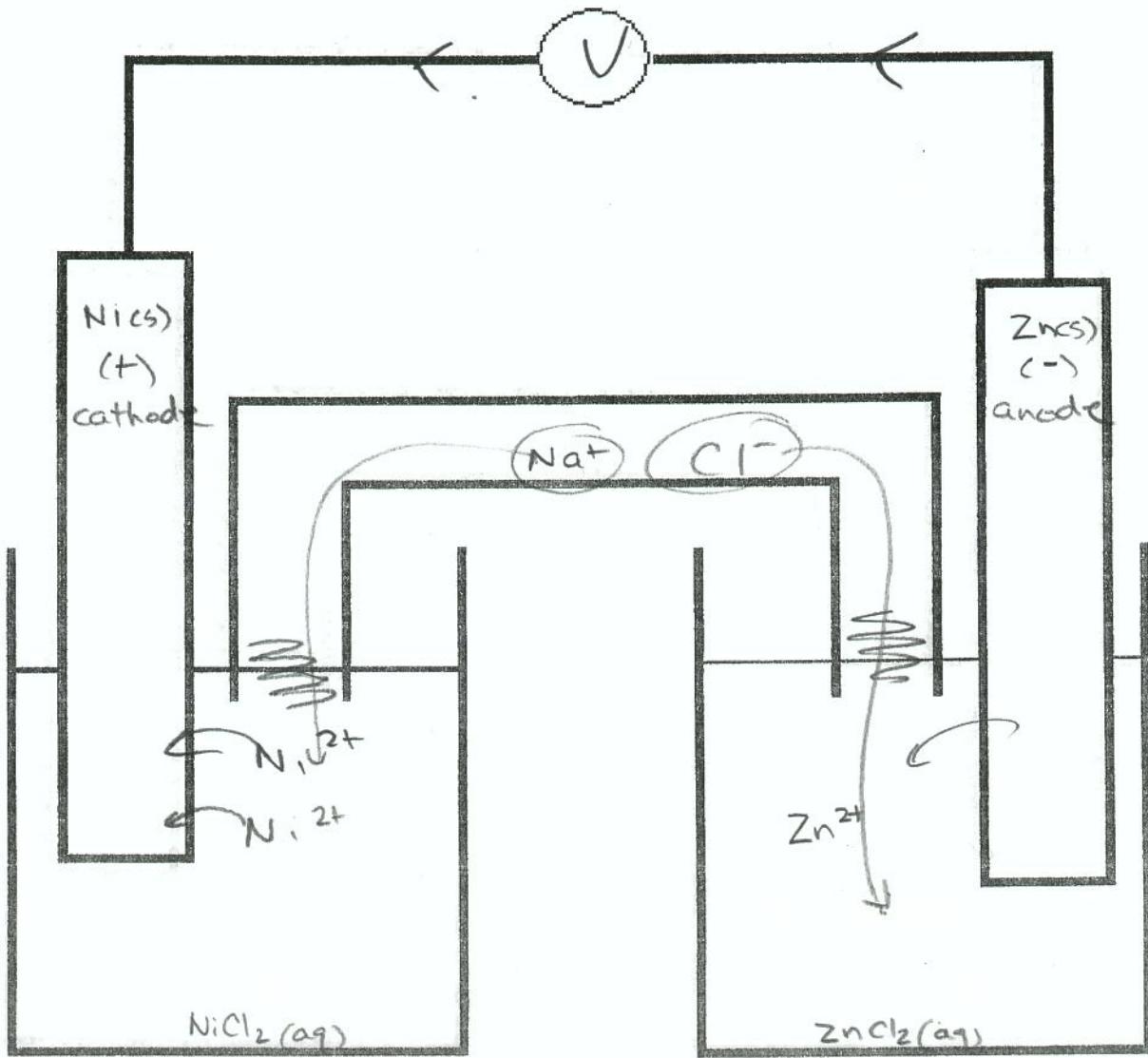
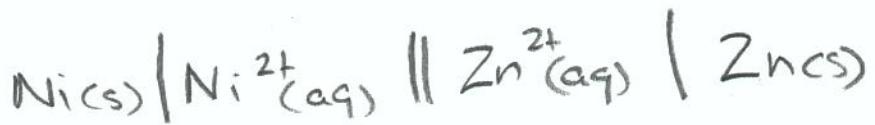
$$\Sigma^\circ = 0.40\text{ V}$$



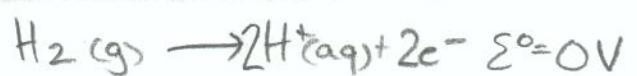
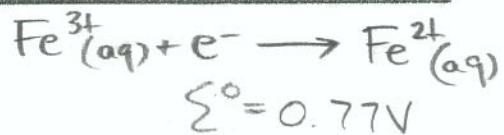
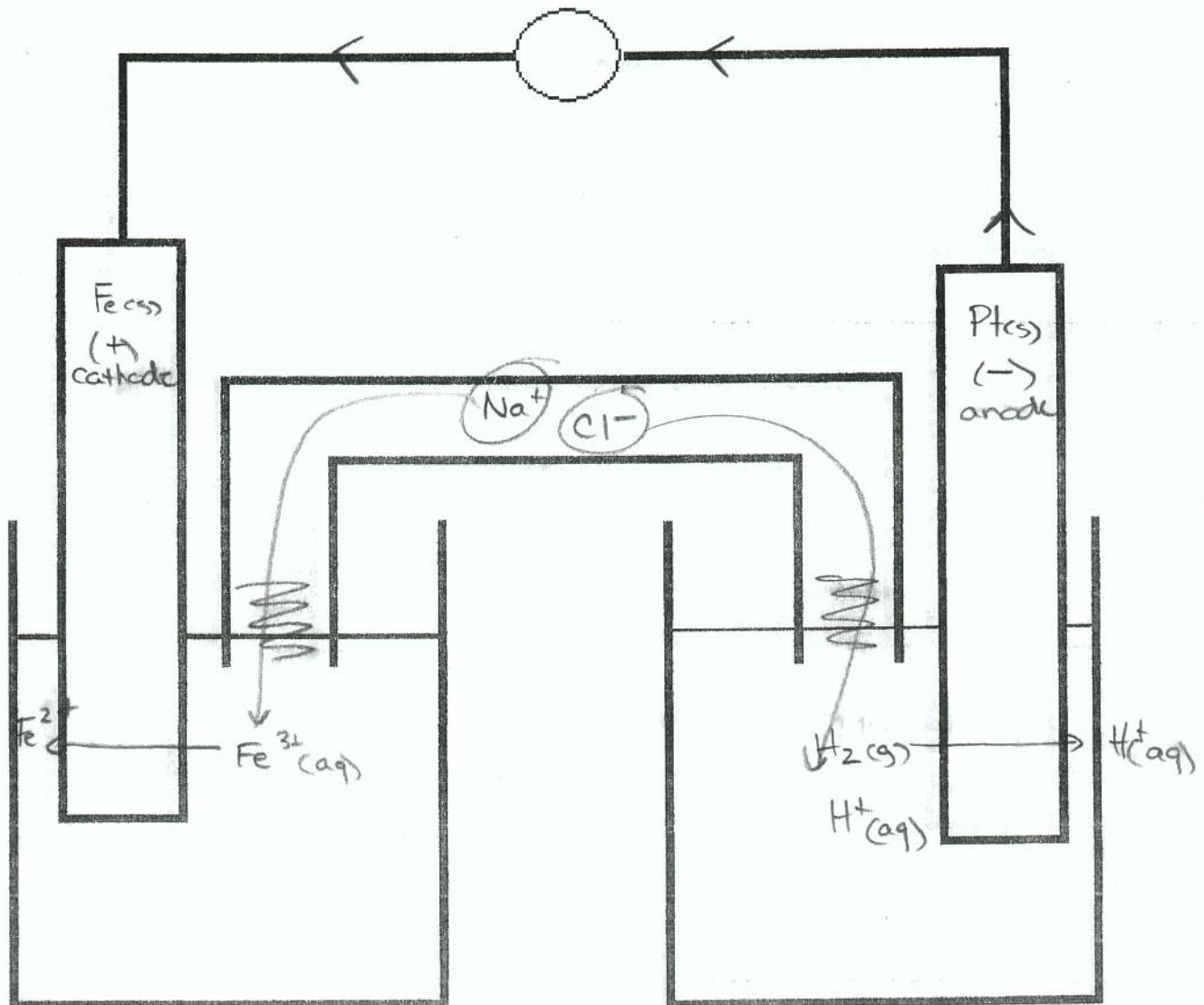
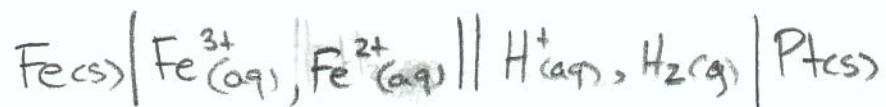
$$\Sigma^\circ = 0.83\text{ V}$$



$$\begin{aligned}\Sigma^\circ_{\text{cell}} &= \Sigma \text{ red} + \Sigma \text{ ox} \\ &= 0.40 + 0.83 = 1.23\text{ V}\end{aligned}$$



$$\Sigma^{\circ}_{\text{cell}} = \Sigma^{\circ}_{\text{red}} + \Sigma^{\circ}_{\text{ox}} \\ = -0.26 + 0.76 = 0.50 \text{ V}$$



$$\sum^\circ_{\text{cell}} = \sum^\circ_{\text{red}} + \sum^\circ_{\text{oxid}}$$

$$= 0.77 + 0$$

$$= 0.77\text{V}$$